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Electrolyte Dissolved in  
Water. 250.0 ml-a  
mark 10.0 g  $\text{Na}_2\text{SO}_4$  or  
water  $7.04 \times 10^{-2}$  moles  
 $\text{Na}^+$  or  $14.1 \times 10^{-2}$   
moles  $\text{Na}^+$  and  $10.0 \times 10^{-2}$   
 $7.04 \times 10^{-2}$  moles  $\text{SO}_4^{2-}$   
Critical Thinking  
Questions 6. Verify that  
when 10.0 g of sodium  
sulfate dissolves in  
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reacts with 4.60 grams  
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reacts with 4.60 grams  
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chloride is dissolved in  
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flask which is then  
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metal ions is found to  
be 0.43 moles ...

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Determine the aluminum ion concentration and the chloride ion concentration in 0.125 M  $\text{AlCl}_3$ . 2. Determine the chromium ion concentration and the sulfate ion concentration in a 300 mL solution that contains 5.00 g of chromium(III) sulfate. 3.

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